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179-183 Book 2 Chapter 9 Aromatic  
Hydrocarbons Chapter 22 Hydrocarbons  
You are at chapter 22!! Let's goooo!!!! First  
of all, a hydrocarbon is the simplest  
organic compound and it only contains  
carbon and hydrogen. For this, we can use

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our background knowledge to know that since carbon has four valence electrons, it can form four covalent bonds. Also, since hydrogen only has one, it can only make one bond.

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atom. Therefore, the simplest hydrocarbon molecule consists of a carbon atom bonded to four hydrogen atoms,  $\text{CH}_4$ . This substance is called methane and is an excellent fuel and the main component of natural gas. 698 Chapter 22

Hydrocarbons Figure 22-1 Carbon is a nonmetal element located in group 4A of the periodic table. Elements in this

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You ' ll Learn You will compare the structures and properties of alkanes, alkenes, and alkynes. You will recognize and compare the properties of structural isomers and stereoisomers. You will describe how useful hydrocarbons are obtained from natural sources. Chapter 22 Hydrocarbons

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Chapter 22 Hydrocarbon Compounds 241  
position of their atoms a double bond at least one substituent on each carbon of the double bond cis-2-butene symmetrical superimposed true It is a carbon with four different atoms or groups attached. They contain an asymmetric carbon atom.

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a controlled process by which

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hydrocarbons are broken down or rearranged into smaller, more useful molecules. Related Flashcards Chapter 22 Organic Chemistry Vocabulary

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786 Chapter 22 • Substituted Hydrocarbons and Their Reactions  
Section 22.12.1 Figure 22.1 All of these items contain at least one of the functional groups that you will study in this chapter. For example, the fruit and flowers have sweet-smelling aromas that are due to ester molecules. Objectives Define functional group, and give examples.

## Chapter 22: Substituted Hydrocarbons and Their Reactions

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Chapter 22: Hydrocarbons. STUDY. PLAY. Petroleum. Complex mixture of carbon compounds produced by heat and pressure acting on the remains of once-living organisms buried deep beneath Earth's surface. Vitalism.

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This unit looks at aliphatic and aromatic hydrocarbons, including alkyl substituted hydrocarbons and structural and geometric isomers. There is also a brief introduction to functional groups. Chapter 22 - Hydrocarbon Compounds

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position of their atoms a double bond at least one substituent on each carbon of the double bond cis-2-butene symmetrical superimposed true It is a carbon with four different atoms or groups attached. They contain an asymmetric carbon atom.  
trans-2-butene  $\text{CH}_2\text{CH}_3$

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Chapter 22: Organic Compounds. 22.1 Hydrocarbons; 22.2 Unsaturated Hydrocarbons; 22.3 Functional Groups; 22.4 Alcohols and Ethers; 22.5 Amines and Alkaloids; 22.6 Carbonyl Compounds;

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22.7 Polymers; Chapter 23: The Nutrients of Life. 23.1 Biomolecules Are Molecules Produced and Used by Organisms; 23.2 Carbohydrates Give Structure and Energy

## 22.2 Unsaturated Hydrocarbons | Conceptual Academy

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ChemFacts. • The primary source of hydrocarbons is petroleum. • About 75 million barrels of petroleum are pumped out of the Earth each day. • Hydrocarbons are used as fuels and are

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the raw materials for products such as plastics, synthetic fibers, solvents, and industrial chemicals.

## Chapter 21: Hydrocarbons

Chapter 22 Hydrocarbons 700 Chapter 22 Hydrocarbons Figure 22-4 This taxi in Bangkok, Thailand is one of many cars, trucks, and buses that have been modified to burn propane. Burning propane benefits the environment because it produces less air pollution than does gasoline. In Table

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Chapter 22 Hydrocarbons Chapter 22 Hydrocarbon Compounds239 SECTION 22.2 UNSATURATED

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HYDROCARBONS (pages 702 – 703)

This section explains the difference between unsaturated and saturated Page 4/30

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