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Rate Law

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AP Chemistry Chapter 14 Answers – Zumdahl 14.39 a. $\text{HClO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_3\text{O}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. Only the forward reaction is indicated since HClO_4 is a strong acid and is basically 100% dissociated in water. For acids, the dissociation reaction is commonly written without water as a reactant. The common abbreviation for this

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AP Chemistry Chapter 14 Answers – Zumdahl 14.41 $\text{pH} = -\log[\text{H}^+]$; $\text{pOH} = -\log[\text{OH}^-]$; At 25 oC, $\text{pH} + \text{pOH} = 14.00$; For Exercise 13.37: a. $\text{pH} = -\log[\text{H}^+] = -\log(1.0 \times 10^{-7}) = 7.00$; $\text{pOH} = 14.00 - \text{pH} = 14.00 - 7.00 = 7.00$ b. $\text{pH} = -\log(8.3 \times 10^{-16}) = 15.08$; $\text{pOH} = 14.00 - 15.08 = -1.08$ c. $\text{pH} = -\log(12) = -1.08$; $\text{pOH} = 14.08 - (-1.08) = 15.08$ d.

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AP Chemistry Chapter 14. Chemical Kinetics - 1 - Change in time Change in the concentration of B Average rate with respect to B = Chapter 14. Chemical Kinetics 14.1 Factors that Affect Reaction Rates

- Chemical kinetics = the study of how fast chemical reactions occur.
- Factors which affect rates of reactions:

Chapter 14. Chemical Kinetics

AP CHEMISTRY – MCHS Page | 9 14.5 Temperature and Rate (EOCQ ' S 47 - 59 page 620)

- Most reactions speed up as temperature increases.
- We can illustrate this with chemiluminescent Cyalume® light sticks.
- A chemiluminescent reaction produces light.
- Two light sticks are placed in water, one at room temperature and one in ice.

Chapter 14. Chemical Kinetics

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Answer: Using Equation 14.14, we find $[\text{NO}_2] = 1.00 \times 10^{-3} \text{ M}$ Practice Exercise Solution (continued) 18. Sample Exercise 14.9 Determining the Half-Life of a First-Order Reaction The reaction of $\text{C}_4\text{H}_9\text{Cl}$ with water is a first-order reaction.

AP Chemistry Chapter 14 Sample Exercises

Answer: carbon, by 18.1°C Calculate the work done for the expansion of oxygen gas from 2.5 liters to 3.7 liters against a pressure of 720 torr. Answer: -1.14 liter · atm 20.0 mL of pure water at 285 K is mixed with 48 mL of water at 315 K. What is the final temperature of the mixture in kelvins?

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